ABSTRACT

The reaction of 2-pyridine aldehyde di-n-propylacetal catalyzed by haloethenes in contrast to those catalyzed by Lewis acids, N-halocompounds etc., has received only little attention. Aliphatic acetals give ether and alcohols as the major products, while aromatic homocyclic acetals yield esters and ethers as the main products. This feature induced the authors to take up the title investigation. Haloethenes are synthetically very useful reagents and vary widely in their acceptor synthon character and reactivity, hence their application in the present work.

KEYWORDS: 2-Pyridine aldehyde di-n-propylacetal, trichloroethylene

INTRODUCTION

Acetals play a vital role in bioorganic research in exploring, antimalarial, antiviral, antibacterial, anti-inflammatory, antitumor and anticancer activities. The action of heterocyclic aromatic acetals by Lewis acids, the rearrangements of aromatic acetals over solid acids, action of aromatic and heteroaromatic acetals by halo compounds, are reported in literature.

Acetal derivatives of aldehyde are valuable in synthesis either as intermediates or as protecting groups. It is known that acetals are susceptible to addition, oxidation, reduction, rearrangement, condensation and hydrolysis in presence of catalysts.

The use of titanium chloride and stannic chloride in organic reactions is of relatively recent origin. Titanium tetrachloride and its alkoxy derivatives have been used in many reactions.

Antony investigated the effect of ring substituent on the mechanism of rearrangement of aromatic acetals by the action of acid chloride. The reaction conducted in 1,2-dichloroethane has been postulated to pass through carboxylation intermediate which on subsequent alkoxy coordination yields esters. Apart from the metal halides and alkoxides, non-metallic compounds such as boron trifluoride, iodine and its inorganic derivatives have also been used extensively by various researchers.

Alphonse has investigated the action of various Lewis acids on aromatic acetals and proposed mechanism based on carboxylation intermediate.

The reactions of aliphatic acetals catalyzed by solid acids have been extensively studied by various researchers and in most cases synthetically important α, β-unsaturated ethers have been obtained as the major products. The unsaturated ether was shown to be formed by the elimination of alkoxy group and the removal of a proton from the β-carbon of acetal. The elimination may occur either in a concerted or stepwise manner depending on the nature of the catalyst. The results of the action of haloethenes on 2-Pyridine aldehyde di-n-propylacetal in acetonitrile medium are reported in the present work.

MATERIALS

Substrate: 2-Pyridine aldehyde di-n-propylacetal was prepared and its purity was checked spectroscopically.

Solvent: The acetonitrile was purified by standard method and used as the solvent.

Reagents: BDH samples of trichloroethylene and tetrachloroethylene were bought and used for the reactions.

EXPERIMENTAL SECTION

1. Acetal preparation.

2. Pyridine aldehyde di-n-propylacetal

3. 53.5gm (0.5mol) of freshly vacuum distilled 2-pyridine aldehyde and 89gm (1.2mol) of distilled n-propyl alcohol were taken in a 500ml round-bottomed flask fitted with a Dean-Stark apparatus carrying a reflux condenser attached to a calcium chloride guard tube. 0.05g of p-toluene sulfonic acid and 80ml of pure dry benzene were added to the solution and the mixture was refluxed for 6 hours.

The flask was cooled to room temperature and the contents were washed with 1M sodium bicarbonate solution and then with water. The solvent was dried over potassium carbonate. After evaporation of the solvent, the crude acetal was distilled under reduced pressure. Pure acetal was collected at 186°C (10mm of Hg) and the yield was 65%.

\[ \eta = 1.464 \text{ at } 34^\circ C \]  
\[ \text{IR: } 1020-1140 \text{ cm}^{-1} (\text{C-O-C}) \]  
\[ \text{PMR: } 0.9, 1.3-1.7, 3.5 (\text{n-propyl}), 5.45 (1H, s, pyridyl-CH), 6.5-7.5 (4H, m, pyridine-H) \]

2. Action of haloethene compounds on 2-pyridine aldehyde di-n-propylacetal

a) Reaction of tetrachloroethylene with 2-pyridine aldehyde di-n-propylacetal 5 ml of 2-pyridine aldehyde di-n-propylacetal in 10 ml of acetonitrile was taken in a 250 ml conical flask, 3.5 g of tetrachloroethylene was dissolved in 10 ml of acetonitrile and was added drop wise to the same flask, fitted with a magnetic stirrer. The temperature was kept at -20°C. The reaction mixture was stirred and the stirring was continued for 1 hour. The reaction mixture was washed with water and the product was extracted with diethyl ether. The resulting reaction mixture was spotted at the TLC. The product was separated by column chromatography and was identified by IR, PMR, CMR and mass spectrum to be the ester.

\[ \text{b) Reaction of trichloroethylene with 2-pyridine aldehyde di-n-propylacetal} \]

5 ml of 2-pyridine aldehyde di-n-propylacetal in 10 ml of acetonitrile was taken in a 250 ml conical flask, 2.7 g of trichloroethylene was dissolved in 10 ml of acetonitrile and was added drop wise to the same flask, fitted with a magnetic stirrer. The temperature was kept at -20°C. The reaction mixture was stirred and the stirring was continued for 1 hour. The reaction mixture was washed with water and the product was extracted with diethyl ether. The resulting reaction mixture was spotted at the TLC. The product was separated by column chromatography and was identified by IR, PMR, CMR and mass spectrum to be the aldehyde.

RESULTS AND DISCUSSION

The action of haloethene compounds on 2-pyridine aldehyde di-n-propylacetal The acetal (1) contains the benzal carbon atom which is surrounded by one H atom and the other three bulky groups namely benzene ring and the two butoxy groups. Thus the acetal requires steric relief. Thus the butoxy oxygen atom with two lone pairs is longing to attack any acceptor synthon. This is the driving force for the attack of the alkoxy oxygen on the acceptor synthon.

The action of tetrachloroethylene and trichloroethylene on the 2-pyridine aldehyde di-n-propylacetal gave the corresponding ester and aldehyde...
respectively. The formation of ester indicates that this reaction may pass through the mechanism as shown in scheme–1.

Tetrachloroethylene (2) is an alkene which is ready to draw nucleophile towards it, since the four chlorine atoms are having electron withdrawing inductive effect (-I effect) thus the acetal (1) makes use of its lone pair on the alkoxy oxygen and forms oxonium ion (3). Each of the chlorine atoms of the oxonium ion intermediate (3) is posing the steric hindrance to the R group of the alkoxy oxygen atom of the oxonium ion. The steric hindrance experienced by the R group is the driving force for the R to get cleaved as RH abstracting the methine H atom. Thus the intermediate (4) is resulted. The reagent tetrachloroethylene which has acted as the catalyst is relieved when the ester (5) is formed as the product.

The trichloroethylene (6) is susceptible for nucleophilic attack like (3) at the acetal (1) resulting in the oxonium ion (7). In this case since the acceptor synthon is only trichloroethylene, the R group attached to the oxonium oxygen is not so much sterically hindered as in the case of (3). Hence the R group found in the other alkoxy group is cleaved resulting in aldehyde (8) as shown in scheme–2.

CONCLUSION

The reactions of the 2-pyridine aldehyde di-n-propylacetal with tetrachloroethylene and trichloroethylene were studied at -20°C. The products formed and the proposed mechanisms followed are given. The reagents tetrachloroethylene and trichloroethylene were found to yield the products ester and aldehyde respectively.

The formation of ester is explained by the mechanism in scheme–1 while the formation of aldehyde is explained by the mechanism in scheme–2. Just as the tetrachloroethylene and the trichloroethylene, many more acceptor synthons can be used to react with the acetals and such reactions can be run. Many more aromatic nuclei like furan, pyrrole, thiophene, pyridine etc. other than the benzene nucleus can be used in the synthesis of the acetals. The same reagents used in the present study can also be treated with aliphatic acetais, heteroaromatic acetals with different hetero atoms such as N, S and O.

REFERENCES